

# University of Pretoria Yearbook 2022

## General chemistry 117 (CMY 117)

<b>Qualification</b>	Undergraduate
<b>Faculty</b>	Faculty of Natural and Agricultural Sciences
<b>Module credits</b>	16.00
<b>NQF Level</b>	05
<b>Programmes</b>	BEEd (Senior Phase and Further Education and Training Teaching)
	BSc (Computer Science)
	BDietetics
	BSc (Meteorology)
	BSc (Applied Mathematics)
	BSc (Biochemistry)
	BSc (Biological Sciences)
	BSc (Biotechnology)
	BSc (Chemistry)
	BSc (Culinary Science)
	BSc (Ecology)
	BSc (Engineering and Environmental Geology)
	BSc (Entomology)
	BSc (Food Science)
	BSc (Genetics)
	BSc (Geography and Environmental Science)
	BSc (Geology)
	BSc (Human Genetics)
	BSc (Human Physiology)
	BSc (Human Physiology, Genetics and Psychology)
	BSc (Mathematics)
	BSc (Medical Sciences)
	BSc (Microbiology)
	BSc (Nutrition)



BSc (Physics)

BSc (Plant Science)

BSc (Zoology)

BScAgric (Agricultural Economics and Agribusiness Management)

BScAgric (Animal Science)

BScAgric (Applied Plant and Soil Sciences)

BScAgric (Plant Pathology)

**Service modules**

Faculty of Engineering, Built Environment and Information Technology

Faculty of Education

Faculty of Health Sciences

Faculty of Veterinary Science

**Prerequisites**

A candidate must have Mathematics for at least 60% and 60% for Physical Sciences.

**Contact time**

1 practical per week, 4 lectures per week

**Language of tuition**

Module is presented in English

**Department**

Chemistry

**Period of presentation**

Semester 1

**Module content**

General introduction to inorganic, analytical and physical chemistry. Atomic structure and periodicity. Molecular structure and chemical bonding using the VSEOR model. Nomenclature of inorganic ions and compounds. Classification of reactions: precipitation, acid-base, redox reactions and gas-forming reactions. Mole concept and stoichiometric calculations concerning chemical formulas and chemical reactions. Principles of reactivity: energy and chemical reactions. Physical behaviour gases, liquids, solids and solutions and the role of intermolecular forces. Rate of reactions: Introduction to chemical kinetics.

The regulations and rules for the degrees published here are subject to change and may be amended after the publication of this information.

The [General Academic Regulations \(G Regulations\)](#) and [General Student Rules](#) apply to all faculties and registered students of the University, as well as all prospective students who have accepted an offer of a place at the University of Pretoria. On registering for a programme, the student bears the responsibility of ensuring that they familiarise themselves with the General Academic Regulations applicable to their registration, as well as the relevant faculty-specific and programme-specific regulations and information as stipulated in the relevant yearbook. Ignorance concerning these regulations will not be accepted as an excuse for any transgression, or basis for an exception to any of the aforementioned regulations.